



IDENTIFICATION AND STRUCTURAL CHARACTERIZATION OF $\text{Ca}(\text{Pt}_x\text{Cu}_{1-x})_3\text{O}_4$

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ABSTRACT

We describe a new compound, $\text{Ca}(\text{Pt}_x\text{Cu}_{1-x})_3\text{O}_4$ ($0.2 < x < 0.3$), which is only the second-known quaternary compound of an alkaline earth metal with platinum, copper, and oxygen. The composition and structure were determined with an electron probe using wave-dispersive spectrometry (WDS), electron diffraction, and powder X-ray diffraction. The present material crystallizes in the Pt bronze structure, a cubic cell with $a = 5.6270(2)$ Å, and space group $Pm\bar{3}n$ (no. 223). The Cu and Pt atoms are randomly distributed onto the same position, comprising nonintersecting chains in each of the three coordinate directions. The material is an n-type semiconductor. © 2000 Elsevier Science Ltd

KEYWORDS: A. oxides, A. semiconductors, C. X-ray diffraction, D. crystal structure

INTRODUCTION

Platinum is used extensively as a crucible material in high-temperature chemical processes. One strong motivation for understanding its reactivity with alkaline earth (AE) metals in the presence of Cu arises from its application in the processing of high- T_c superconducting

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cuprates, all of which contain at least one AE metal. The present material was first detected in the residue left on a Pt crucible where a mixture of CaCuO_2 and CuO was melted at 1050°C . The chemical composition was determined by wave-dispersive spectrometry (WDS) using an electron microprobe, and the approximate composition was used to prepare samples that were essentially phase-pure.

To our knowledge, several ternary AE–Pt–O and Pt–Cu–O compounds have been identified, such as CaPt_2O_4 [1], $\text{Ca}_2\text{Pt}_3\text{O}_8$ [2], M_4PtO_6 (where $\text{M} = \text{Ba}$ or Ca) [3–5], $\text{Cu}_{0.8}\text{Pt}_{0.2}\text{O}$, and Cu_3PtO_6 [6], but only one quaternary AE–Pt–Cu–O compound, namely, $\text{Ba}_4\text{Pt}_{1-x}\text{Cu}_{2-x}\text{O}_{9-z}$ [7].

EXPERIMENTAL

Bulk samples of $\text{Ca}(\text{Pt}_x\text{Cu}_{1-x})_3\text{O}_4$ were prepared by calcining mixtures containing CaO , CuO , and Pt metal powders at 1050°C in flowing oxygen. The samples were furnace-cooled in oxygen and ground after every 10 h of heat treatment, and parts of the samples were analyzed by laboratory X-ray diffraction (XRD). While the diffraction patterns were found to become unchanged after 10 h of heat treatment, an additional 20 h was imposed to ensure complete reaction.

Chemical analysis was carried out by WDS [8] with a JEOL 8600 electron probe, which had been calibrated using compounds with known compositions (Cu and O calibrated on Cu_2O , Ca on $\text{CaMgSi}_2\text{O}_6$, Pt on Pt metal) and had an accuracy of $\pm 3\%$. Electron diffraction was conducted with a JEOL 2000FX transmission electron microscope (TEM), which operated at 200 kV and was equipped with an energy-dispersive X-ray (EDX) analyzer. The samples were prepared by crushing, and small particles were dispersed on a holed carbon film (Ni grid).

Laboratory powder XRD was carried out under step-scan mode ($3 \text{ sec}/0.02^\circ$) with a Mac Science/MXP diffractometer using $\text{Cu K}\alpha$ radiation. Initial synchrotron XRD data were collected at an X-ray wavelength of 0.49935 \AA in a capillary of 0.65 mm diameter at beamline X3B1 of the National Synchrotron Light Source at Brookhaven National Laboratory. The diffractometer is based on a $\text{Si}(111)$ monochromator and $\text{Ge}(111)$ analyzer, with a commercial $\text{NaI}(\text{Tl})$ scintillation detector. The capillary geometry was chosen to eliminate the possibility of preferred orientation before the cubic lattice was known. Once the structure was determined from these data, we collected additional spectra of several samples for refinement, using samples side-drifted into a flat-plate sample holder, at an X-ray wavelength of 1.15 \AA .

Electrical characterization was carried out on bar-shaped bulk samples. The resistivity was determined by the standard four-point method, while the carrier type was determined by Hall effect measurement under a field of 0.6 Tesla at room temperature.

RESULTS AND DISCUSSION

Backscattering microscopy and WDS analyses showed that the original residue contained predominantly one quaternary oxide, whose cation compositions showed a fairly constant $(\text{Pt} + \text{Cu})/\text{Ca}$ ratio of ~ 3.0 ($\pm 3\%$) but some spreading in the Pt/Cu ratio near 0.4 ($\pm 15\%$). Bulk samples with the cation composition of $\text{Ca}:\text{Pt}:\text{Cu} = 1:3x:3(1-x)$, where x varies from 0.1 to 0.4 with a minimum increment of 0.03, were thus prepared and subjected to laboratory XRD analysis (Fig. 1). Samples with x in the range of 0.2 and 0.3 gave essentially the same set of Bragg reflections (Table 1), which also appeared as the major phase in the pattern of the original residue.

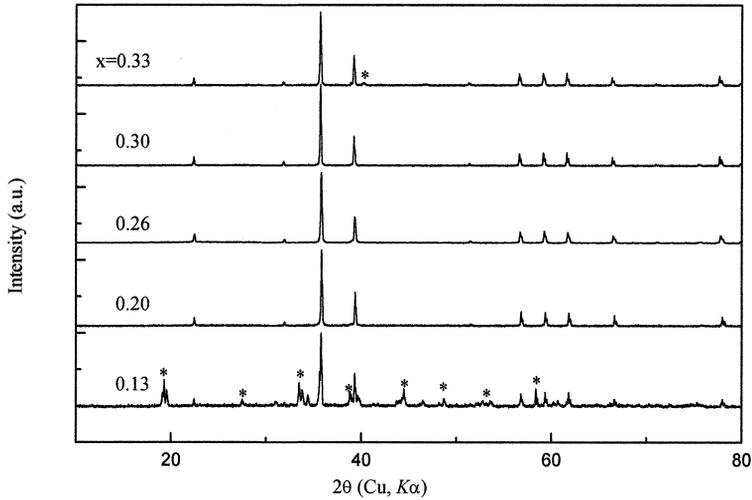


FIG. 1

The Cu $K\alpha$ X-ray diffraction patterns for calcined samples of nominal compositions $\text{Ca}(\text{Pt}_x\text{Cu}_{1-x})_3\text{O}_4$. Peaks marked with an asterisk (*) are due to impurity phases.

The synchrotron data set for the $x = 0.33$ sample was indexed by noting that the strongest lines were regularly spaced in $\sin^2\theta$, suggesting a cubic lattice. The reflection conditions— $(00l)$ l even; (hhl) l even; $(0kl)$ and (hkl) never extinct except for (014) —indicate either the $P\bar{4}3n$ or $Pm\bar{3}n$ space group. This symmetry was also observed in the single-crystal electron diffraction patterns. The size of the cell implied that the atoms must be in Wyckoff positions of high symmetry, and trial and error established that the occupations must be a mixture of Pt and Cu in $6c$ $(1/4, 0, 1/2)$, Ca in $2a$ $(0, 0, 0)$, and O in $8e$ $(1/4, 1/4, 1/4)$. This assignment exactly matches the $\text{Ca}/(\text{Cu},\text{Pt})$ ratio determined from WDS. It is also in agreement with density measurement, which, carried out based on the principle of Archimedes, showed a density of 8.2 ± 0.5 and hence suggested $Z = 2$. No extra diffraction lines or spots due to superstructure were detected in any X-ray or EDX measurements, indicating random mixing between Pt and Cu. The absence of the (014) peak is consistent with the restriction of atoms to Wyckoff positions a , c , and e .

Structure refinement was carried out by the Rietveld method, based on the synchrotron XRD pattern of an $x = 0.33$ sample. The result, for space group $Pm\bar{3}n$, lattice parameter $a = 5.6270(2)$ Å, is shown in Figure 2. In this spectrum, we can observe weak peaks from two different phases described as Ca_4PtO_6 [4,5] and broad, asymmetric lines from an fcc phase with lattice parameter 3.87 Å, assumed to be an alloy of Cu and Pt. Excluding the regions of impurity peaks, the Rietveld fit was $\chi^2 = 7.46$; this rather bad fit can be explained by the difficulties of ascertaining lineshape for inadequately sampled particle statistics with very sharp peaks. We were unable to determine reliable thermal parameters, probably also due to inadequate sampling. The Bragg R-factor of 0.030 (defined as

$$\frac{\sum_k |I_k - I_{c_k}|}{\sum_k I_k}$$

TABLE 1
The Cu K α X-ray Diffraction Data for Nominal
Composition CaPt_{0.8}Cu_{2.2}O₄

<i>hkl</i>	<i>d</i> (Å)	<i>I</i> / <i>I</i> ₀
110	3.983	11
200	2.816	4
210	2.519	100
211	2.300	39
310	1.781	2
222	1.626	14
320	1.562	15
321	1.506	12
400	1.408	9
411	1.327	1
330		
420	1.260	1
421	1.229	9
332	1.201	2
510	1.105	<1
431		
432	1.046	8
520		
521	1.029	2
440	0.996	3
442	0.939	<1
600		
610	0.926	1
532	0.914	2
611		

for experimental observations I_k and calculated intensities I_{c_k}) is more appropriate for the present case of fully resolved peaks. The fractional Pt occupancy x refines to 0.22 ± 0.05 . The spectrum is also consistent with space group $P\bar{4}3n$, representing a particular (unlikely) buckling of the four oxygens surrounding each (Pt,Cu), with a maximum allowed oxygen displacement of 0.3 Å.

In order to explore the limit of Pt:Cu composition, we collected a synchrotron X-ray powder pattern from an $x = 0.13$ sample. That sample had larger amounts of both Ca₄PtO₆ phases and a trace of CuO along with the Ca(Pt,Cu)₃O₄ phase described here. The other phases were so strong that it was not possible to obtain a Rietveld refinement, but three peaks were observed without significant contamination, and their ratio implied a Pt occupancy x in the range of 0.22 ± 0.08 . The lattice parameter was 5.622(1)Å.

Figure 3 shows the overall view of the crystal structure of the present material. Infinite chains of (Cu,Pt) extend, parallel to the edges of the unit cube. They are tied together by oxygen ions to form interlocking cages, at the center of which are Ca ions. Each (Cu,Pt) ion is surrounded by four oxygen ions in square planar coordination, while each Ca ion has eight oxygen ions as its closest neighbors. No unusual bond distances or coordinates can be seen

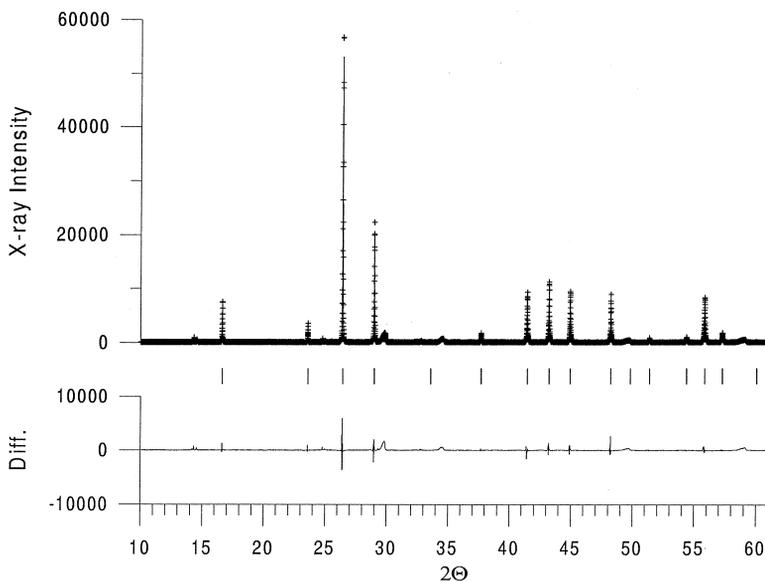


FIG. 2

Rietveld refinement of a sample of nominal composition $\text{CaPtCu}_2\text{O}_4$.

(Table 2). The Ca–O distance is 2.437 Å, while the (Cu/Pt)–O distance is 1.989 Å. The sp^2 hybridization of the oxygen atoms accounts for the planar-triangular bonding to (Cu,Pt) ions.

This material is isostructural to the well-known platinum bronzes, which have a general formula of $\text{M}_y\text{Pt}_3\text{O}_4$ [9]. Known M metals include Na, Mg, Ni, and Cd, with y being no greater than 1.0. By taking the normal valence of the M element, such as +1 for Na, +2 for Mg, and Cd, and +2 to about +4 for Ni, and multiplying it by the known y values, one

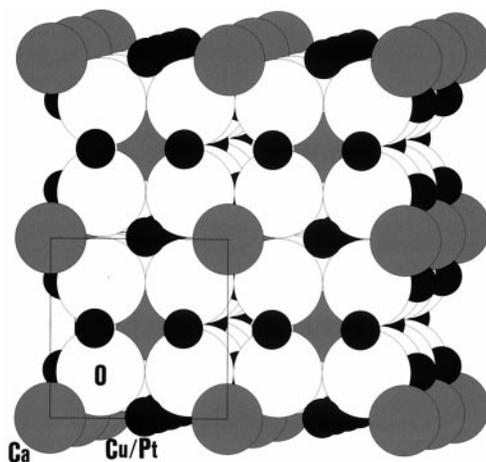


FIG. 3

The proposed structure of $\text{Ca}(\text{Pt}_x\text{Cu}_{(1-x)})_3\text{O}_4$.

TABLE 2
Interatomic Distances (Å)

(Cu,Pt)–O	1.989
(Cu,Pt)–(Cu,Pt)	2.813
(Cu,Pt)–Ca	3.146
O–O	2.813
Ca–O	2.437

readily finds out that the total valence contributed by the M ions is no greater than +1 per formula in the known Pt bronzes. Cahen *et al.* [1] attempted to prepare $\text{Ca}_x\text{Pt}_3\text{O}_4$ but ended up with CaPt_2O_4 , which has a structure similar to the bronzes, except that the Pt chain along one of the axes is systematically missing. It appears that there existed a minimum valence, no less than 7/3, for Pt ions to be stable in the square-planar coordination with oxygen. When the minimum valence was not fulfilled, stability was regained by reducing Pt ions in each unit cell, as in the case of CaPt_2O_4 , or by replacing Pt with another element that could take a lower valence than Pt. The latter has been demonstrated previously for CdPd_3O_4 [10] and CaPd_3O_4 [11], and now in the present compound. The observed Pt:Cu ratio in the present compound coincides well with $\text{Cu}_{0.8}\text{Pt}_{0.2}\text{O}$ [6], where the Cu/Pt ion is also in square-planar coordination with oxygen ions.

Resistivity measurements of samples with $x = 0.2$ to ~ 0.3 gave data characteristic of a semiconductor, showing an increase in resistivity with decreasing temperature. The room-temperature resistivities of the polycrystalline samples were found to decrease with increasing Pt content, ranging from 0.8 $\Omega\text{-cm}$ for $x = 0.3$ to 9.9 $\Omega\text{-cm}$ for $x = 0.2$. Hall effect measurements indicated electrons (*i.e.*, *n*-type) to be the major carrier. This is in great contrast to metallic $\text{Ni}_x\text{Pt}_3\text{O}_4$ and *p*-type semiconducting CaPd_3O_4 .

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REFERENCES

1. D. Cahen, J.A. Ibers, and M.H. Mueller, *Inorg. Chem.* **13**, 110 (1974).
2. X. Turrillas, C. Laviron, H. Vincent, J. Pannetier, and J.C. Joubert, *J. Solid State Chem.* **67**, 297 (1987).
3. S.J. Schneider and C.L. McDaniel, *J. Am. Ceram. Soc.* **52**, 518 (1969).
4. R. Czaya, *Z. Anorg. Allg. Chem.* **375**, 61 (1970); JCPDS file no. 23-860.

5. JCPDS file no. 23-862.
6. O. Muller and R. Roy, *J. Less-Common Met.* **19**, 209 (1969).
7. JCPDS file 40-302.
8. E.P. Bertin, *Introduction to X-ray Spectrometric Analysis*, Plenum Press, New York (1978), Chaps. 1-4, 6, 7, 9.
9. J. Waser and E.D. McClanahan, Jr., *J. Chem. Phys.* **19**, 413 (1951).
10. O. Muller and R. Roy, *Advan. Chem. Ser.* **98**, 28 (1971).
11. R.C. Wnuk, T.R. Touw, and B. Post, *IBM J. Res. Devel.* **8**, 185 (1964).